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Understanding acid-base
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students in chemistry, biochemistry, biology, environmental and earth sciences. Solving chemical problems, be it in education or in real life, often requires the

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understanding of the acid-base equilibria behind them. Based on many years of teaching experience, Heike Kahlert and Fritz Scholz present a powerful tool to meet such

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challenges. They provide a simple guide to the fundamentals and applications of acid-base diagrams, avoiding complex mathematics. This textbook is richly illustrated and

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This book will give students a thorough grounding in pH and associated equilibria, material absolutely fundamental to the understanding of many

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aspects of chemistry. It is, in addition, a fresh and modern approach to a topic all too often taught in an out-moded way. This book uses new theoretical developments which have

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led to more generalized approaches to equilibrium problems; these approaches are often simpler than the approximations which they replace. Acid-base problems are readily

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addressed in terms of the proton condition, a convenient amalgam of the mass and charge constraints of the chemical system considered. The graphical

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approach of Bjerrum, Hagg, and Sillen is used to illustrate the orders of magnitude of the concentrations of the various species involved in chemical equilibria.

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Based on these concentrations, the proton condition can usually be simplified, often leading directly to the value of the pH. In the description of acid-base titrations a

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general master equation is developed. It provides a continuous and complete description of the entire titration curve, which can then be used for computer-based comparison with

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experimental data.

Graphical estimates of the steepness of titration curves are also developed, from which the practicality of a given titration can be

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anticipated. Activity effects are described in detail, including their effect on titration curves. The discussion emphasizes the distinction between equilibrium

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constants and
electrometric pH
measurements, which are
subject to activity
corrections, and balance
equations and
spectroscopic pH

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measurements, which are not. Finally, an entire chapter is devoted to what the pH meter measures, and to the experimental and theoretical uncertainties involved.

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Selected Proceedings from

the CAL 85 Symposium

Acid-Base Diagrams

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Biochemistry

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Parasitic infection can greatly interfere with a soldier's ability to complete his mission. The presence of parasites in a soldier's system can not only interfere with his ability to*

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function, but also can make him susceptible to certain diseases. Since soldiers may serve in most areas of the world, you must be able to identify parasites that are found in the various parts of the globe. In

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your job as a medical laboratory specialist, you will perform a variety of test procedures on samples taken from humans. Some of these samples will include feces and tissue scrapings used in the diagnosis

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and treatment of parasitic infection. Therefore, you must be knowledgeable in several areas of parasitology. The knowledge you will need is reflected in the two subcourses you are about to study.

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Subcourses Parasitology I and Parasitology II address areas of particular importance in parasitology. The whole purpose of clinical laboratory procedures is to provide the clinician doing diagnostic work with specific

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information needed to round out his picture of the disorders he has observed in the patient.

Clinical bacteriology can contribute its part by supplying data about the microscopic life involved and the susceptibility

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of such life to particular drugs. To identify bacterial growth, you must take certain steps that will enable you, through a process of elimination, to choose the microscopic form that fits the findings you have obtained.

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Steps that are often essential include: 1. Observing the type of growth when first isolated on culture media. 2 Making a microscopic examination on stained material from an isolated culture of that colony.

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3. Performing various tests to obtain a list of the characteristics of the organism.

4. Making a complete identification of the organism.

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mathematical skills as a Medical Laboratory Specialist. The emphasis is upon computations related to solutions and their concentrations. If you feel that you need a more basic review of mathematics before taking this

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subcourse, you should request Subcourse Basic Mathematics, which covers addition, subtraction, multiplication, and division of whole numbers; decimals, and fractions; and conversions to and from the

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metric system. In the process of achieving and maintaining proficiency in your military occupational specialty (MOS), you will be learning concepts and performing tasks that are based on important chemical

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principles. As you become more proficient with these principles, you may reach the point where you will not need to give them much conscious thought.

Meanwhile, however, you should study this subcourse to gain a

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working knowledge of the fundamental principles of chemistry. Subcourse Clinical Chemistry I, provides you with a background in the laboratory basics of clinical chemistry. Laboratory safety; collection,

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preservation, and shipment of specimens; measurement of weights and volumes; introduction to quality control; and introduction to organic chemistry are presented in this subcourse.

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builds upon previous editions with more enhanced content that deals with principles and techniques of quantitative analysis with more examples of analytical techniques drawn from areas such as clinical

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textbook. Each chapter contains Overview, Topical Outline, Skills, and Common Mistakes, which are all keyed to the textbook for easy cross reference. The Overview section summarizes the content of the chapter and

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includes a comprehensive listing of terms, a summary of general concepts, and a list of numerical exercises, while the Topical Outline provides the subtopic heads that carry the corresponding chapter and

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section numbers as they appear in the textbook. The Fill-in, Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides

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developed exercises to apply the new concepts in the chapter to particular examples. The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry,

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while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter. After briefly dealing with an overview

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of chemistry, this book goes on exploring the concept of matter, energy, measurement, problem solving, atom, periodic table, and chemical bonding. These topics are followed by discussions on writing names

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and formulas of compounds; chemical formulas and the mole; chemical reactions; calculations based on equations; gases; and the properties of a liquid. The remaining chapters examine the

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solutions; acids; bases; salts; oxidation-reduction reactions; electrochemistry; chemical kinetics and equilibrium; and nuclear, organic, and biological chemistry. This study guide will be of great value to chemistry

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Interest in comparative acid-base physiology has considerably

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grown during last decades even in the absence of major technical or conceptual advances. This is firstly because it has become clear that the extracellular acid-base state reflects the performance of many exchange functions at the organism level:

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respiration and ventilation of the gas exchange surfaces, metabolism, iono-and osmoregulation. Such functions are much influenced by ambient conditions, and the measurement of acid base parameters thus provides useful information about the

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organism's responses to environmental challenges. Secondly, many processes at the molecular level are now known to be pH sensitive, and acid-base regulation thus appears to be a major requirement for the functional integrity of cells and

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organisms. How extracellular acid-base balance can be maintained in a wide variety of animals living in different conditions is the subject of this book. The approach is comparative and environmental throughout. All body fluids share

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similar buffer proper ties, and common physicochemical principles apply to any acid base system. However, in accord with differing designs and con straints along animal evolution, varying effector organs and mechanisms are at work to maintain an

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appropriate acid-base state in the organism. Particular emphasis is placed on the fundamental differences between water and air breathers and on the acid-base and respiratory problems arising at the transition from an aquatic to a terrestrial

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Kjeldahl method, and molarity and moles where EDTA and water hardness calculations are concerned. New discussions of the Ag/AgCl and combination pH electrodes have been added, while the discussion of ion-selective

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Analysis

Basic material; Useful statistics; Stoichiometry; Acid-base chemistry; General discussion; Strong acids and bases; Weak monoprotic acids; Weak polyprotic acids; The titration weak acids; The titration of weak

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acids; Weak bases; The control of solubility; Complex formation with a ligand; Formation of a weak acid with the anion of the precipitate; Absorption spectrophotometry; Basic relationships and instrumentation; Some

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applications of absorption spectrophotometry; Potentiometry; Basic considerations; Potentiometric titrations; Physical constants and activities; Ion-selective electrodes; Separations not involving precipitation; Transfer

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of solutes between two immiscible liquid phases using separatory funnels; Multistage liquid-liquid extractions; Liquid-liquid chromatography (LLC); Gas-liquid chromatography (GLC); Separations by sorption chromatography.

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facilitate learning, and the accompanying Interactive Examples in OWLv2 have been redesigned to better parallel the problem-solving approach in the narrative. New Capstone Problems have been added to a number of chapters. Important Notice: Media

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